

Chemical equations worksheet pdf

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Balancing Chemical Equations

Balance the equations below:

- 1) $\underline{\quad}$ N₂ + $\underline{\quad}$ H₂ \rightarrow $\underline{\quad}$ NH₃
- 2) $\underline{\quad}$ KClO₃ \rightarrow $\underline{\quad}$ KCl + $\underline{\quad}$ O₂
- 3) $\underline{\quad}$ NaCl + $\underline{\quad}$ F₂ \rightarrow $\underline{\quad}$ NaF + $\underline{\quad}$ Cl₂
- 4) $\underline{\quad}$ H₂ + $\underline{\quad}$ O₂ \rightarrow $\underline{\quad}$ H₂O
- 5) $\underline{\quad}$ Pb(OH)₂ + $\underline{\quad}$ HCl \rightarrow $\underline{\quad}$ H₂O + $\underline{\quad}$ PbCl₂
- 6) $\underline{\quad}$ AlBr₃ + $\underline{\quad}$ K₂SO₄ \rightarrow $\underline{\quad}$ KBr + $\underline{\quad}$ Al₂(SO₄)₃
- 7) $\underline{\quad}$ CH₄ + $\underline{\quad}$ O₂ \rightarrow $\underline{\quad}$ CO₂ + $\underline{\quad}$ H₂O
- 8) $\underline{\quad}$ C₃H₈ + $\underline{\quad}$ O₂ \rightarrow $\underline{\quad}$ CO₂ + $\underline{\quad}$ H₂O
- 9) $\underline{\quad}$ C₆H₁₆ + $\underline{\quad}$ O₂ \rightarrow $\underline{\quad}$ CO₂ + $\underline{\quad}$ H₂O
- 10) $\underline{\quad}$ FeCl₃ + $\underline{\quad}$ NaOH \rightarrow $\underline{\quad}$ Fe(OH)₃ + $\underline{\quad}$ NaCl
- 11) $\underline{\quad}$ P + $\underline{\quad}$ O₂ \rightarrow $\underline{\quad}$ P₂O₅
- 12) $\underline{\quad}$ Na + $\underline{\quad}$ H₂O \rightarrow $\underline{\quad}$ NaOH + $\underline{\quad}$ H₂
- 13) $\underline{\quad}$ Ag₂O \rightarrow $\underline{\quad}$ Ag + $\underline{\quad}$ O₂
- 14) $\underline{\quad}$ S₈ + $\underline{\quad}$ O₂ \rightarrow $\underline{\quad}$ SO₃
- 15) $\underline{\quad}$ CO₂ + $\underline{\quad}$ H₂O \rightarrow $\underline{\quad}$ C₆H₁₂O₆ + $\underline{\quad}$ O₂
- 16) $\underline{\quad}$ K + $\underline{\quad}$ MgBr \rightarrow $\underline{\quad}$ KBr + $\underline{\quad}$ Mg
- 17) $\underline{\quad}$ HCl + $\underline{\quad}$ CaCO₃ \rightarrow $\underline{\quad}$ CaCl₂ + $\underline{\quad}$ H₂O + $\underline{\quad}$ CO₂
- 18) $\underline{\quad}$ HNO₃ + $\underline{\quad}$ NaHCO₃ \rightarrow $\underline{\quad}$ NaNO₃ + $\underline{\quad}$ H₂O + $\underline{\quad}$ CO₂
- 19) $\underline{\quad}$ H₂O + $\underline{\quad}$ O₂ \rightarrow $\underline{\quad}$ H₂O₂
- 20) $\underline{\quad}$ NaBr + $\underline{\quad}$ CaF₂ \rightarrow $\underline{\quad}$ NaF + $\underline{\quad}$ CaBr₂
- 21) $\underline{\quad}$ H₂SO₄ + $\underline{\quad}$ NaNO₂ \rightarrow $\underline{\quad}$ HNO₂ + $\underline{\quad}$ Na₂SO₄

Balancing Equations Practice Worksheet

Balance the following equations:

- 1) $\underline{\quad}$ NaNO₃ + $\underline{\quad}$ PbO \rightarrow $\underline{\quad}$ Pb(NO₃)₂ + $\underline{\quad}$ Na₂O
- 2) $\underline{\quad}$ AgI + $\underline{\quad}$ Fe₂(CO₃)₃ \rightarrow $\underline{\quad}$ FeI₃ + $\underline{\quad}$ Ag₂CO₃
- 3) $\underline{\quad}$ C₂H₄O₂ + $\underline{\quad}$ O₂ \rightarrow $\underline{\quad}$ CO₂ + $\underline{\quad}$ H₂O
- 4) $\underline{\quad}$ ZnSO₄ + $\underline{\quad}$ Li₂CO₃ \rightarrow $\underline{\quad}$ ZnCO₃ + $\underline{\quad}$ Li₂SO₄
- 5) $\underline{\quad}$ V₂O₅ + $\underline{\quad}$ CaS \rightarrow $\underline{\quad}$ CaO + $\underline{\quad}$ V₂S₅
- 6) $\underline{\quad}$ Mn(NO₂)₂ + $\underline{\quad}$ BeCl₂ \rightarrow $\underline{\quad}$ Be(NO₂)₂ + $\underline{\quad}$ MnCl₂
- 7) $\underline{\quad}$ AgBr + $\underline{\quad}$ GaPO₄ \rightarrow $\underline{\quad}$ Ag₃PO₄ + $\underline{\quad}$ GaBr₃
- 8) $\underline{\quad}$ H₂SO₄ + $\underline{\quad}$ B(OH)₃ \rightarrow $\underline{\quad}$ B₂(SO₄)₃ + $\underline{\quad}$ H₂O
- 9) $\underline{\quad}$ S₈ + $\underline{\quad}$ O₂ \rightarrow $\underline{\quad}$ SO₂
- 10) $\underline{\quad}$ Fe + $\underline{\quad}$ AgNO₃ \rightarrow $\underline{\quad}$ Fe(NO₃)₂ + $\underline{\quad}$ Ag

Solutions for the Balancing Equations Practice Worksheet

- 1) 2 NaNO₃ + PbO \rightarrow Pb(NO₃)₂ + Na₂O
- 2) 6 AgI + Fe₂(CO₃)₃ \rightarrow 2 FeI₃ + 3 Ag₂CO₃
- 3) C₂H₄O₂ + 2 O₂ \rightarrow 2 CO₂ + 2 H₂O
- 4) ZnSO₄ + Li₂CO₃ \rightarrow ZnCO₃ + Li₂SO₄

Balancing Equations Worksheet

- 1) $\text{H}_3\text{PO}_4 + \text{KOH} \rightarrow \text{K}_3\text{PO}_4 + \text{H}_2\text{O}$
- 2) $\text{K} + \text{B}_2\text{O}_3 \rightarrow \text{K}_2\text{O} + \text{B}$
- 3) $\text{HCl} + \text{NaOH} \rightarrow \text{NaCl} + \text{H}_2\text{O}$
- 4) $\text{Na} + \text{NaNO}_3 \rightarrow \text{Na}_2\text{O} + \text{N}_2$
- 5) $\text{C} + \text{S}_8 \rightarrow \text{CS}_2$
- 6) $\text{Na} + \text{O}_2 \rightarrow \text{Na}_2\text{O}$
- 7) $\text{N}_2 + \text{O}_2 \rightarrow \text{N}_2\text{O}_5$
- 8) $\text{H}_3\text{PO}_4 + \text{Mg(OH)}_2 \rightarrow \text{Mg}_3(\text{PO}_4)_2 + \text{H}_2\text{O}$
- 9) $\text{NaOH} + \text{H}_2\text{CO}_3 \rightarrow \text{Na}_2\text{CO}_3 + \text{H}_2\text{O}$
- 10) $\text{KOH} + \text{HBr} \rightarrow \text{KBr} + \text{H}_2\text{O}$
- 11) $\text{Na} + \text{O}_2 \rightarrow \text{Na}_2\text{O}$
- 12) $\text{Al(OH)}_3 + \text{H}_2\text{O}_2 \rightarrow \text{Al}_2(\text{CO}_3)_3 + \text{H}_2\text{O}$
- 13) $\text{Al} + \text{S}_8 \rightarrow \text{Al}_2\text{S}_3$
- 14) $\text{Cs} + \text{N}_2 \rightarrow \text{Cs}_2\text{N}$
- 15) $\text{Mg} + \text{Cl}_2 \rightarrow \text{MgCl}_2$
- 16) $\text{Rb} + \text{RbNO}_3 \rightarrow \text{Rb}_2\text{O} + \text{N}_2$
- 17) $\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$
- 18) $\text{N}_2 + \text{H}_2 \rightarrow \text{NH}_3$
- 19) $\text{C}_{10}\text{H}_{16} + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$
- 20) $\text{Al(OH)}_3 + \text{HBr} \rightarrow \text{AlBr}_3 + \text{H}_2\text{O}$
- 21) $\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_3 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$
- 22) $\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$
- 23) $\text{Li} + \text{AlCl}_3 \rightarrow \text{LiCl} + \text{Al}$
- 24) $\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$
- 25) $\text{NH}_3\text{OH} + \text{H}_3\text{PO}_4 \rightarrow (\text{NH}_4)_2\text{PO}_4 + \text{H}_2\text{O}$
- 26) $\text{Rb} + \text{P} \rightarrow \text{Rb}_3\text{P}$
- 27) $\text{CH}_4 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$
- 28) $\text{Al(OH)}_3 + \text{H}_2\text{SO}_4 \rightarrow \text{Al}_2(\text{SO}_4)_3 + \text{H}_2\text{O}$
- 29) $\text{Na} + \text{Cl}_2 \rightarrow \text{NaCl}$
- 30) $\text{Rb} + \text{S}_8 \rightarrow \text{Rb}_2\text{S}$
- 31) $\text{H}_3\text{PO}_4 + \text{Ca(OH)}_2 \rightarrow \text{Ca}_3(\text{PO}_4)_2 + \text{H}_2\text{O}$
- 32) $\text{NH}_3 + \text{HCl} \rightarrow \text{NH}_4\text{Cl}$
- 33) $\text{Li} + \text{H}_2\text{O} \rightarrow \text{LiOH} + \text{H}_2$
- 34) $\text{Ca}_3(\text{PO}_4)_2 + \text{SiO}_2 + \text{C} \rightarrow \text{CaSiO}_3 + \text{CO} + \text{P}$
- 35) $\text{NH}_3 + \text{O}_2 \rightarrow \text{N}_2 + \text{H}_2\text{O}$
- 36) $\text{FeS}_2 + \text{O}_2 \rightarrow \text{Fe}_2\text{O}_3 + \text{SO}_2$
- 37) $\text{C} + \text{SO}_2 \rightarrow \text{CS}_2 + \text{CO}$

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Balancing Equations Worksheet

- 1) $\text{Na}_3\text{PO}_4 + \text{KOH} \rightarrow \text{NaOH} + \text{K}_3\text{PO}_4$
- 2) $\text{MgF}_2 + \text{Li}_2\text{CO}_3 \rightarrow \text{MgCO}_3 + \text{LiF}$
- 3) $\text{P}_4 + \text{O}_2 \rightarrow \text{P}_2\text{O}_5$
- 4) $\text{RbNO}_3 + \text{BeF}_2 \rightarrow \text{Be}(\text{NO}_3)_2 + \text{RbF}$
- 5) $\text{AgNO}_3 + \text{Cu} \rightarrow \text{Cu}(\text{NO}_3)_2 + \text{Ag}$
- 6) $\text{Cl}_2 + \text{Br}_2 \rightarrow \text{ClBr} + \text{F}_2$
- 7) $\text{HCN} + \text{CuSO}_4 \rightarrow \text{H}_2\text{SO}_4 + \text{Cu}(\text{CN})_2$
- 8) $\text{GaF}_3 + \text{Cs} \rightarrow \text{ClF} + \text{Ga}$
- 9) $\text{BaS} + \text{PbF}_2 \rightarrow \text{BaF}_2 + \text{PbS}$
- 10) $\text{N}_2 + \text{H}_2 \rightarrow \text{NH}_3$
- 11) $\text{NaF} + \text{Br}_2 \rightarrow \text{NaBr} + \text{F}_2$
- 12) $\text{Pb(OH)}_2 + \text{HCl} \rightarrow \text{H}_2\text{O} + \text{PbCl}_2$
- 13) $\text{AlBr}_3 + \text{K}_2\text{SO}_4 \rightarrow \text{KBr} + \text{Al}(\text{SO}_4)_3$
- 14) $\text{CH}_4 + \text{O}_2 \rightarrow \text{CO} + \text{H}_2\text{O}$
- 15) $\text{Na}_3\text{PO}_4 + \text{CaCl}_2 \rightarrow \text{NaCl} + \text{Ca}_3(\text{PO}_4)_2$
- 16) $\text{K} + \text{Cl}_2 \rightarrow \text{KCl}$
- 17) $\text{Al} + \text{HCl} \rightarrow \text{H}_2 + \text{AlCl}_3$
- 18) $\text{N}_2 + \text{F}_2 \rightarrow \text{NF}_3$
- 19) $\text{SO}_2 + \text{LiSe} \rightarrow \text{SSe}_2 + \text{Li}_2\text{O}$
- 20) $\text{NH}_3 + \text{H}_2\text{SO}_4 \rightarrow (\text{NH}_4)_2\text{SO}_4$

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In order to continue enjoying our site, we ask that you confirm your identity as a human. Thank you very much for your cooperation. A balanced chemical equation gives the number and type of atoms participating in a reaction, the reactants, products, and direction of the reaction. Balancing an unbalanced equation is mostly a matter of making certain mass and charge are balanced on the reactants and products side of the reaction arrow. This is a collection of printable worksheets to practice balancing equations. The printable worksheets are provided in pdf format with separate answer keys. Balancing Chemical Equations - Worksheet #1Balancing Chemical Equations - Answers #1Balancing Chemical Equations - Worksheet #2Balancing Chemical Equations - Answers #2Balancing Chemical Equations - Worksheet #3Balancing Chemical Equations - Answers #3Balancing Equations - Worksheet #4Balancing Equations - Answers Key #4 I also offer printable worksheets for balancing equations on my personal site. The printables are also available as PDF files: Balancing Equation Practice Sheet [answer sheet]Another Equation Worksheet [answer sheet]Yet Another Printable Worksheet [answer key] You may also wish to review the step-by-step tutorial on how to balance a chemical equation. Another way to practice balancing equations is by taking a quiz. Coefficients in Balanced Equations QuizBalance Chemical Equations Quiz These scaffolded Balancing Chemical Equations Cornell Doodle Notes combine two effective note-taking strategies and can be used to introduce, teach, or review the concept of balancing chemical equations and the Law of Conservation of Mass. These notes review chemical formulas including coefficients and subscripts, differentiate between word, skeletal, and balanced equations, use two non-chemistry analogies to help students visualize what a balanced equation means, give steps for balancing an equPage 2

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