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Balancing Chemical Equations

Balance the equations below:

- 1) $\text{N}_2 + \text{H}_2 \rightarrow \text{NH}_3$
- 2) $\text{KClO}_3 \rightarrow \text{KCl} + \text{O}_2$
- 3) $\text{NaCl} + \text{F}_2 \rightarrow \text{NaF} + \text{Cl}_2$
- 4) $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$
- 5) $\text{Pb}(\text{OH})_2 + \text{HCl} \rightarrow \text{H}_2\text{O} + \text{PbCl}_2$
- 6) $\text{AlBr}_3 + \text{K}_2\text{SO}_4 \rightarrow \text{KBr} + \text{Al}_2(\text{SO}_4)_3$
- 7) $\text{CH}_4 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$
- 8) $\text{C}_3\text{H}_8 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$
- 9) $\text{C}_8\text{H}_{18} + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$
- 10) $\text{FeCl}_3 + \text{NaOH} \rightarrow \text{Fe}(\text{OH})_3 + \text{NaCl}$
- 11) $\text{P} + \text{O}_2 \rightarrow \text{P}_2\text{O}_5$
- 12) $\text{Na} + \text{H}_2\text{O} \rightarrow \text{NaOH} + \text{H}_2$
- 13) $\text{Ag}_2\text{O} \rightarrow \text{Ag} + \text{O}_2$
- 14) $\text{S}_8 + \text{O}_2 \rightarrow \text{SO}_3$
- 15) $\text{CO}_2 + \text{H}_2\text{O} \rightarrow \text{C}_6\text{H}_{12}\text{O}_6 + \text{O}_2$
- 16) $\text{K} + \text{MgBr} \rightarrow \text{KBr} + \text{Mg}$
- 17) $\text{HCl} + \text{CaCO}_3 \rightarrow \text{CaCl}_2 + \text{H}_2\text{O} + \text{CO}_2$
- 18) $\text{HNO}_3 + \text{NaHCO}_3 \rightarrow \text{NaNO}_3 + \text{H}_2\text{O} + \text{CO}_2$
- 19) $\text{H}_2\text{O} + \text{O}_2 \rightarrow \text{H}_2\text{O}_2$
- 20) $\text{NaBr} + \text{CaF}_2 \rightarrow \text{NaF} + \text{CaBr}_2$
- 21) $\text{H}_2\text{SO}_4 + \text{NaNO}_2 \rightarrow \text{HNO}_2 + \text{Na}_2\text{SO}_4$

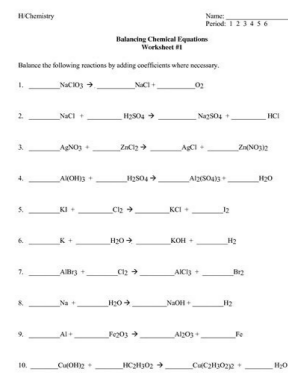
Balancing Equations Practice Worksheet

Balance the following equations:

- 1) $\text{NaNO}_3 + \text{PbO} \rightarrow \text{Pb}(\text{NO}_3)_2 + \text{Na}_2\text{O}$
- 2) $\text{AgI} + \text{Fe}_2(\text{CO}_3)_3 \rightarrow \text{FeI}_3 + \text{Ag}_2\text{CO}_3$
- 3) $\text{C}_2\text{H}_4\text{O}_2 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$
- 4) $\text{ZnSO}_4 + \text{Li}_2\text{CO}_3 \rightarrow \text{ZnCO}_3 + \text{Li}_2\text{SO}_4$
- 5) $\text{V}_2\text{O}_5 + \text{CaS} \rightarrow \text{CaO} + \text{V}_2\text{S}_5$
- 6) $\text{Mn}(\text{NO}_2)_2 + \text{BeCl}_2 \rightarrow \text{Be}(\text{NO}_2)_2 + \text{MnCl}_2$
- 7) $\text{AgBr} + \text{GaPO}_4 \rightarrow \text{Ag}_3\text{PO}_4 + \text{GaBr}_3$
- 8) $\text{H}_2\text{SO}_4 + \text{B}(\text{OH})_3 \rightarrow \text{B}_2(\text{SO}_4)_3 + \text{H}_2\text{O}$
- 9) $\text{S}_8 + \text{O}_2 \rightarrow \text{SO}_2$
- 10) $\text{Fe} + \text{AgNO}_3 \rightarrow \text{Fe}(\text{NO}_3)_2 + \text{Ag}$

Solutions for the Balancing Equations Practice Worksheet

- 1) $2 \text{NaNO}_3 + \text{PbO} \rightarrow \text{Pb}(\text{NO}_3)_2 + \text{Na}_2\text{O}$
- 2) $6 \text{AgI} + \text{Fe}_2(\text{CO}_3)_3 \rightarrow 2 \text{FeI}_3 + 3 \text{Ag}_2\text{CO}_3$
- 3) $\text{C}_2\text{H}_4\text{O}_2 + 2 \text{O}_2 \rightarrow 2 \text{CO}_2 + 2 \text{H}_2\text{O}$
- 4) $\text{ZnSO}_4 + \text{Li}_2\text{CO}_3 \rightarrow \text{ZnCO}_3 + \text{Li}_2\text{SO}_4$



Balancing Equations Worksheet

- 1) $\text{H}_3\text{PO}_4 + \text{KOH} \rightarrow \text{K}_3\text{PO}_4 + \text{H}_2\text{O}$
- 2) $\text{K} + \text{B}_2\text{O}_3 \rightarrow \text{K}_2\text{O} + \text{B}$
- 3) $\text{HCl} + \text{NaOH} \rightarrow \text{NaCl} + \text{H}_2\text{O}$
- 4) $\text{Na} + \text{NaNO}_3 \rightarrow \text{Na}_2\text{O} + \text{N}_2$
- 5) $\text{C} + \text{S}_8 \rightarrow \text{CS}_2$
- 6) $\text{Na} + \text{O}_2 \rightarrow \text{Na}_2\text{O}$
- 7) $\text{N}_2 + \text{O}_2 \rightarrow \text{N}_2\text{O}_5$
- 8) $\text{H}_3\text{PO}_4 + \text{Mg(OH)}_2 \rightarrow \text{Mg}_2(\text{PO}_4)_3 + \text{H}_2\text{O}$
- 9) $\text{NaOH} + \text{H}_2\text{CO}_3 \rightarrow \text{Na}_2\text{CO}_3 + \text{H}_2\text{O}$
- 10) $\text{KOH} + \text{HBr} \rightarrow \text{KBr} + \text{H}_2\text{O}$
- 11) $\text{Na} + \text{O}_2 \rightarrow \text{Na}_2\text{O}$
- 12) $\text{Al(OH)}_3 + \text{H}_2\text{CO}_3 \rightarrow \text{Al}_2(\text{CO}_3)_3 + \text{H}_2\text{O}$
- 13) $\text{Al} + \text{S}_8 \rightarrow \text{Al}_2\text{S}_3$
- 14) $\text{Cs} + \text{N}_2 \rightarrow \text{Cs}_3\text{N}$
- 15) $\text{Mg} + \text{Cl}_2 \rightarrow \text{MgCl}_2$
- 16) $\text{Rb} + \text{RbNO}_3 \rightarrow \text{Rb}_2\text{O} + \text{N}_2$
- 17) $\text{C}_6\text{H}_6 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$
- 18) $\text{N}_2 + \text{H}_2 \rightarrow \text{NH}_3$
- 19) $\text{C}_{12}\text{H}_{22} + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$
- 20) $\text{Al(OH)}_3 + \text{HBr} \rightarrow \text{AlBr}_3 + \text{H}_2\text{O}$
- 21) $\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_3 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$
- 22) $\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$
- 23) $\text{Li} + \text{AlCl}_3 \rightarrow \text{LiCl} + \text{Al}$
- 24) $\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$
- 25) $\text{NH}_4\text{OH} + \text{H}_3\text{PO}_4 \rightarrow (\text{NH}_4)_3\text{PO}_4 + \text{H}_2\text{O}$
- 26) $\text{Rb} + \text{P} \rightarrow \text{Rb}_3\text{P}$
- 27) $\text{CH}_4 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$
- 28) $\text{Al(OH)}_3 + \text{H}_2\text{SO}_4 \rightarrow \text{Al}_2(\text{SO}_4)_3 + \text{H}_2\text{O}$
- 29) $\text{Na} + \text{Cl}_2 \rightarrow \text{NaCl}$
- 30) $\text{Rb} + \text{S}_8 \rightarrow \text{Rb}_2\text{S}$
- 31) $\text{H}_3\text{PO}_4 + \text{Ca(OH)}_2 \rightarrow \text{Ca}_3(\text{PO}_4)_2 + \text{H}_2\text{O}$
- 32) $\text{NH}_3 + \text{HCl} \rightarrow \text{NH}_4\text{Cl}$
- 33) $\text{Li} + \text{H}_2\text{O} \rightarrow \text{LiOH} + \text{H}_2$
- 34) $\text{Ca}_3(\text{PO}_4)_2 + \text{SiO}_2 + \text{C} \rightarrow \text{CaSiO}_3 + \text{CO} + \text{C}$
- 35) $\text{NH}_3 + \text{O}_2 \rightarrow \text{N}_2 + \text{H}_2\text{O}$
- 36) $\text{FeS}_2 + \text{O}_2 \rightarrow \text{Fe}_2\text{O}_3 + \text{SO}_2$
- 37) $\text{C} + \text{SO}_2 \rightarrow \text{CS}_2 + \text{CO}$

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Balancing Equations Worksheet

- 1) $\text{Na}_3\text{PO}_4 + \text{KOH} \rightarrow \text{NaOH} + \text{K}_3\text{PO}_4$
- 2) $\text{MgF}_2 + \text{Li}_2\text{CO}_3 \rightarrow \text{MgCO}_3 + \text{LiF}$
- 3) $\text{P}_4 + \text{O}_2 \rightarrow \text{P}_2\text{O}_5$
- 4) $\text{RbNO}_3 + \text{BeF}_2 \rightarrow \text{Be(NO}_3)_2 + \text{RbF}$
- 5) $\text{AgNO}_3 + \text{Cu} \rightarrow \text{Cu(NO}_3)_2 + \text{Ag}$
- 6) $\text{CF}_4 + \text{Br}_2 \rightarrow \text{CBr}_4 + \text{F}_2$
- 7) $\text{HCN} + \text{CuSO}_4 \rightarrow \text{H}_2\text{SO}_4 + \text{Cu(CN)}_2$
- 8) $\text{GaF}_3 + \text{Ca} \rightarrow \text{CaF}_2 + \text{Ga}$
- 9) $\text{BaS} + \text{PF}_5 \rightarrow \text{BaF}_2 + \text{PS}$
- 10) $\text{N}_2 + \text{H}_2 \rightarrow \text{NH}_3$
- 11) $\text{NaF} + \text{Br}_2 \rightarrow \text{NaBr} + \text{F}_2$
- 12) $\text{Pb(OH)}_2 + \text{HCl} \rightarrow \text{H}_2\text{O} + \text{PbCl}_2$
- 13) $\text{AlBr}_3 + \text{K}_2\text{SO}_4 \rightarrow \text{KBr} + \text{Al}_2(\text{SO}_4)_3$
- 14) $\text{CH}_4 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$
- 15) $\text{Na}_3\text{PO}_4 + \text{CaCl}_2 \rightarrow \text{NaCl} + \text{Ca}_3(\text{PO}_4)_2$
- 16) $\text{K} + \text{Cl}_2 \rightarrow \text{KCl}$
- 17) $\text{Al} + \text{HCl} \rightarrow \text{H}_2 + \text{AlCl}_3$
- 18) $\text{N}_2 + \text{F}_2 \rightarrow \text{NF}_3$
- 19) $\text{SO}_2 + \text{Li}_2\text{Se} \rightarrow \text{SSe}_2 + \text{Li}_2\text{O}$
- 20) $\text{NH}_3 + \text{H}_2\text{SO}_4 \rightarrow (\text{NH}_4)_2\text{SO}_4$

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In order to continue enjoying our site, we ask that you confirm your identity as a human. Thank you very much for your cooperation. A balanced chemical equation gives the number and type of atoms participating in a reaction, the reactants, products, and direction of the reaction. Balancing an unbalanced equation is mostly a matter of making certain mass and charge are balanced on the reactants and products side of the reaction arrow. This is a collection of printable worksheets to practice balancing equations. The printable worksheets are provided in pdf format with separate answer keys. Balancing Chemical Equations - Worksheet #1 Balancing Chemical Equations - Answers #1 Balancing Chemical Equations - Worksheet #2 Balancing Chemical Equations - Answers #2 Balancing Chemical Equations - Worksheet #3 Balancing Chemical Equations - Answers #3 Balancing Chemical Equations - Worksheet #4 Balancing Chemical Equations - Answer Key #4 I also offer printable worksheets for balancing equations on my personal site. The printables are also available as PDF files: Balancing Equation Practice Sheet [answer sheet] Another Equation Worksheet [answer sheet] Yet Another Printable Worksheet [answer key] You may also wish to review the step-by-step tutorial on how to balance a chemical equation. Another way to practice balancing equations is by taking a quiz. Coefficients in Balanced Equations Quiz Balance Chemical Equations Quiz These scaffolded Balancing Chemical Equations Cornell Doodle Notes combine two effective note-taking strategies and can be used to introduce, teach, or review the concept of balancing chemical equations and the Law of Conservation of Mass. These notes review chemical formulas including coefficients and subscripts, differentiate between word, skeletal, and balanced equations, use two non-chemistry analogies to help students visualize what a balanced equation means, give steps for balancing an equation Page 2

